**Electron Configuration Worksheet**

**Each sub- shell can hold a specific maximum number of electrons:**

**s= 2 p= 6 d= 10 f=14**

**Example:**

**Barium (Ba) 56 has electrons. Each sub-shell is filled with maximum number of electrons until all 56 electrons are placed in a sub-shell.**

 **Number of Electrons**

 **1s2 2s2 2p6 3s2 3p6 4s2 3d10 4p6 5s2 4d10 5p6 6s2**

 **Subshell**

**Calcium (Ca) has 20 electrons. Fill each sub-shell below with the maximum number of electrons until all are placed in a sub-shell.**

**1s 2s 2p 3s 3p 4s**

**Use the sub-shell configuration below as a guide for finding electron configurations of other elements. Start filling the lowest energy subshell (1s) first.**

**1s 2s 2p 3s 3p 4s 3d 4p 5s 4d 5p 6s 4f 5d 6p 7s 5f 6d**

**\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_**

**Electron Configuration Worksheet**

**Each sub- shell can hold a specific maximum number of electrons:**

**s= 2 p= 6 d= 10 f=14**

**Example:**

**Barium (Ba) 56 has electrons. Each sub-shell is filled with maximum number of electrons until all 56 electrons are placed in a sub-shell.**

 **Number of Electrons**

 **1s2 2s2 2p6 3s2 3p6 4s2 3d10 4p6 5s2 4d10 5p6 6s2**

 **Subshell**

**Calcium (Ca) has 20 electrons. Fill each sub-shell below with the maximum number of electrons until all are placed in a sub-shell.**

**1s 2s 2p 3s 3p 4s**

**Use the sub-shell configuration below as a guide for finding electron configurations of other elements. Start filling the lowest energy subshell (1s) first.**

**1s 2s 2p 3s 3p 4s 3d 4p 5s 4d 5p 6s 4f 5d 6p 7s 5f 6d**